

Session 3: Exam Technique

How do you tackle a question? – 5 minutes



1. Think of **two** ways exam questions give you clues to the answers they are expecting

2. List **three** things you can do when approaching a question in an exam

Common Command Words – 5 mins

Command words are used to instruct the intention of the expected answer. How many do you know?

There are 61 command words used in GCSE! (See Command words document)

Luckily there are only a handful used routinely. Any subject specific commands will be discussed by your teachers:

State

The main points in clear terms

Describe

Give an account by recalling some facts, events or process in an accurate way.

Explain

Students should state the reasons for something happening.

Analyse

Look closely at detail; give reasons why or how something is happening and the effect of this

Compare

Identify similarities and/or differences.

Evaluate

Make a judgement using the information supplied, as well as their knowledge and understanding, to consider evidence for and against when making that judgement

Complete

Finish the task by adding information

Determine

Use given data or information to obtain an answer

Calculate

use numbers given in the question to work out the answer

Clues in the Question – 5 mins

Knowing the command words gives you clues as to how to answer the question. What else do we get that will help?

- The number of marks – an indication of the amount of work you need to demonstrate and the time worth giving the question
- Information before and within the question – data, diagrams, key text that is used to support the question
- Key subject terms – an indication to the topic and where to base your linked ideas

This can be used in summary by BUGging the question...

Box

Box the command words,
What is the question
asking you to do?

Underline

Underline the important parts of
the question but nothing else,
Underline the number of marks
available.

Go Back

Have you responded to the
command words?
Have you addressed the
underlined words?

BUG the Question

The Radcliffe School



Building Bright Futures Together

Give it a go – 10 mins

You will be given three exam questions, each from past papers

1. BUG the first one together – on this PowerPoint
2. Then you can practise on your own

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Question 1 – Together

Q1.(a) **Describe** the model now used for the structure of an atom.

In your answer you should:

- give details of the individual particles that make up an atom
- include the relative masses and relative charges of these particles.

Do **not** include a diagram in your answer.

(6)

Box the command words

Describe

Give an account by recalling some facts, events or process in an accurate way.

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Box the command words.
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Underline Underline the important parts of the question but nothing else, Underline the number of marks available.

Go Back Have you responded to the command words? Have you addressed the underlined words?

Underline the important parts – What do we know?

Describe

Give an account by recalling some facts, events or process in an accurate way.

Atoms have a nucleus and particles orbiting that nucleus. The nucleus is positive. Atoms contain protons which are positive, neutrons which are neutral and electrons which are negative. In the nucleus there are protons and neutrons, both with relative mass of 1. Orbiting are electrons with relative mass of about 1/2000.

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Do **not** include a diagram in your answer.

(6)

Go back and check.

Describe

Give an account by recalling some facts, events or process in an accurate way.

Atoms have a small nucleus and electrons orbiting that nucleus at a relatively large distance from the nucleus. The nucleus is positive. Atoms contain protons which are positive, neutrons which are neutral and electrons which are negative. In the nucleus there are protons and neutrons, both with relative mass of 1. Orbiting are electrons with relative mass of about 1/2000.

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Question 1 – Mark Scheme & Feedback

brief description

contains protons, neutrons and electrons

protons are positive

electrons are negative

neutrons are uncharged

has a nucleus

relative charge

proton +1

electron - 1

neutron 0

relative mass

proton 1

neutron 1

electron (about) 1 / 2000

accept protons and neutrons have the same mass

accept electrons have tiny / negligible mass

zero mass is neutral

more detailed description

protons and neutrons make up the nucleus

electrons orbit the nucleus

electrons are in shells

most of the atom is empty space

nucleus occupies a very small fraction of the volume of the atom

electrons orbit at a relatively large distance from the nucleus

most of the mass of the atom is contained in the nucleus

the nucleus as a whole is positively charged total number of protons in the nucleus equals the total number of electrons orbiting it in an atom

This then goes onto a flash card!

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Now try this with your next questions – 20 mins

Either use the pre-prepared set of questions or subject specific questions to practise BUGging the question

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Finishing Up – 5 mins

Now you have looked at 2 exam questions and assessed...

Produce a flash card for any point that was not mastered and add it to your subject pack.